

Question #1

What is the name of the compound SrO?

Question #2

What is the molar mass for the hydrocarbon $C_{24}H_{37}O_6$

Question #3

What happens to the electrons during a metallic bond?

Question #4

What type of bond forms between two nonmetals share electrons?

Question #5

Balance the following reaction and identify the mole ratio between the two reactants.



Question #6

What are the dominant intermolecular forces present in water?

Question #7

Rank the bulk forces in order of strength from weakest to strongest.

Question #8

What is the formula for copper (IV) sulfate?

Question #9

What kind of a reaction is this?



Question #10

Which molecule has covalent bonding and does not require a double or triple bond?



Question #11

Draw the Lewis dot structure for BrO_3^-

Question #12

What is the VSEPR geometry for carbon tetrachloride?

Question #13

What is the molar mass of $(\text{NH}_4)_2\text{S}$?

Question #14

What is the VSEPR geometry for ammonia?

Question #15

Draw the Lewis Dot structure for nitrogen gas.

Question #16

What kind of a reaction requires O_2 as a reactant?

Question #17

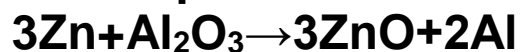
Predict the products and balance the following reaction:
aluminum phosphate plus rubidium nitrite

Question #18

In a 'polar' bond, the elements involved are sharing the electron(s)...

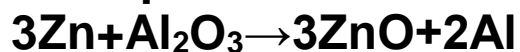
Question #19

If you have 10 mol of Zn, how many mols of ZnO can be produced?



Question #20

If you have 25g of Zn, how many g of ZnO can be produced?



Question #21

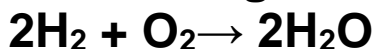
Name the seven diatomic elements.

Question #22

What is the molar mass of $\text{Al}(\text{OH})_3$?

Question #23

If you have 38L of O_2 at STP, how many L of H_2O can you make, assuming your water is gaseous.

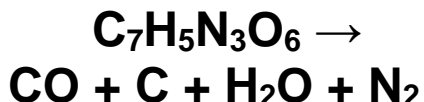


Question #24

Draw the Lewis dot structure for AlH_3

Question #25

What is the mole ratio of TNT to carbon monoxide ?



Question #26

What is the formula of sodium carbonate?

Question #27

What is the molar mass of sodium carbonate?

Question #28

What is the molar mass of iron (III) sulfate

Question #29

How many moles of iron (III) sulfate in 44.5g iron (III) sulfate? The molar mass is 400.1 g/mol.

Question #30

How many grams potassium chloride are in 14.6 moles potassium chloride?
1 mole = 74.5 g

Question #31

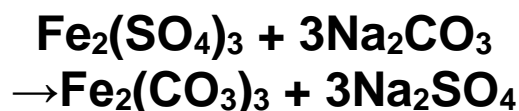
How many grams of iron (III) carbonate are found in 5.46 moles iron (III) carbonate?

Question #33

Aqueous copper(II) bromide reacts with aqueous aluminum chloride. Balance this equation and predict the products.

Question #35

What is the mole ratio of iron(III) sulfate to sodium sulfate?



Question #37

If 15 g oxygen is reacted with hydrogen, then how many moles water will be produced?

Question #39

If 46 g sodium metal reacts completely, how many moles chlorine gas will be required to make sodium chloride?

Question #32

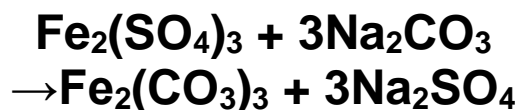
Aqueous copper(II) bromide reacts with aqueous aluminum chloride. What type of reaction is this?

Question #34

How many moles of sodium carbonate are in 10.9 g sodium carbonate? Molar mass is 106 g

Question #36

If 10 moles iron (III) sulfate reacts, how many moles sodium sulfate will form?

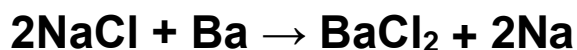


Question #38

What are three of the four indications that a chemical reaction has occurred?

Question #40

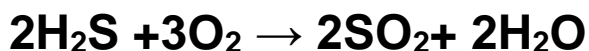
If 100 g sodium chloride reacts completely with barium, then how many grams sodium metal will be obtained?



Question #41

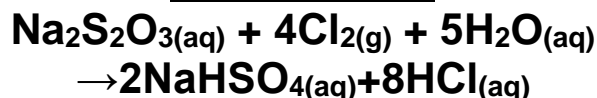
How many liters fluorine gas in 0.87 moles of fluorine gas at STP?

Question #43



How many moles of H_2S are required to form 8.20 moles of SO_2 ?

Question #45



How many moles of H_2O react if 5.24×10^{19} molecules of HCl are formed?

Question #47

Which electrons are involved in bonding?

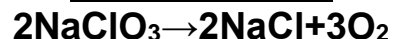
Question #49

How many valence electrons are available in NF_3 ?

Question #42

How many grams of oxygen gas are in 0.69L of oxygen gas at STP?

Question #44



How many molecules of oxygen are produced when 80.0 grams of sodium chloride are produced?

Question #46

Draw the Lewis structure and identify the VSEPR geometry for BCl_3 .

Question #48

How many electrons does Antimony need to gain or lose to have a full valence shell?

Question #50

Identify if the following are ionic or covalent:
 LiF , CH_4 , CH_3OH , NH_3 , MgO

Question #51

List the first ten prefixes used to name covalent molecules

Question #53

Name Ag_2O

Question #55

Name SO_6

Question #57

Draw the Lewis Structure for CH_3OH

Question #59

How many lone pairs does NH_3 have?

Question #52

Name C_2H_6

Question #54

Name $\text{Cu}(\text{NO}_2)_3$

Question #56

Draw the Lewis structure for BaS

Question #58

How many lone pairs does CH_4 have?

Question #60

Is SF_2 polar or non polar?

Question #61

Is SiO_2 polar or non polar?

Question #62

Which is more polar, CHCl_3 or CHBr_3

Question #63

Identify the main IMF present in F_2

Question #64

Identify the main IMF present in HCl

Question #65

Identify the main IMF present in C (graphite)

Question #66

What is the sum of the coefficients when balanced:
 $\text{___ Ag} + \text{___ S} \rightarrow \text{Ag}_2\text{S}$